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Salt is Acidic, Basic, or
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Table*~~~~

10 Amazing Experiments with
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Solutions

Solution: The reaction of
the acid + base gives a
product of salt + water,
which is considered as
neutralization reaction.

Examples: $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl}$
 $+ \text{H}_2\text{O}$. $\text{Mg}(\text{OH})_2 + \text{H}_2\text{CO}_3$
 $\rightarrow \text{MgCO}_3 + 2\text{H}_2\text{O}$. 15. Give
two important uses of
washing soda and baking
soda. Solution:

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chapter 2 Acids, bases and salt

Salts are formed by the combination of acid and base through the neutralization reaction. The acidic and basic nature of salts usually depends on the acid and base from which the salt evolved in neutralization reaction. Example: $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

Acids Bases and Salts |
Properties of Acids, Bases
and Salts

A reaction in which an acid and base react with each other to give a salt and water is termed as neutralization reaction. For Example: (i) $\text{NaOH} + \text{HCl} \rightarrow$

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$\text{NaCl} + \text{H}_2\text{O}$ (ii) $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$ 15. Give two important uses of washing soda and baking soda. Answer Two important uses of washing soda are: → It is used in glass, soap, and paper industries.

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10th Science: Ch 2 Acids,
Bases ...

The presence of the weak acid and salt acts like a "sponge" to soak up excess/additional hydrogen ions or hydroxide ions if an acid or base are added to the solution Determining Acid or Base To determine whether a substance is an acid or a base, count the

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hydrogen on each substance before and after the reaction.

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Class 10 Science Acid, Bases and Salts Mind Map Acids.

Produce $[H^+]$ in H_2O ; Sour taste; Turns blue litmus red; Acts as electrolyte in solution; Chemical

Properties Of Acids Acids + Metal \rightarrow Salt + H_2 e.g. $H_2SO_4 + Zn \rightarrow ZnSO_4 + H_2$

Acids + Metal

Carbonate/Metal hydrogen

Carbonate = Salt + CO_2 e.g. $NaHCO_3 + HCl \rightarrow NaCl + H_2O + CO_2$ Acids + Bases \rightarrow Salt + water

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Science Chapter 2 Acids ...

Salt is an ionic compound that results from the neutralization reaction of acids and bases. Salts are constituted of positively charged ions, known as cations and negatively charged ions, known as anions, which can either be organic or inorganic in nature. These ions are present in a relative amount, thus rendering the nature of the salt neutral.

Acids, Bases, and Salts -
Introduction, Dissociation
...

Most salts do not form pH-neutral solutions. Salts

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such as sodium chloride that can be made by combining a strong acid (HCl) with a strong base (NaOH, KOH) have a neutral pH, but these are exceptions to the general rule that solutions of most salts are mildly acidic or alkaline. Salts of a strong base and a weak acid yield alkaline solutions.

13.3: Finding the pH of weak Acids, Bases, and Salts ...

If _____ ions are found in a solution, the solution is basic in nature. hydrogen . hydroxyl . hydronium Half-n-half Clue : 5. Antacids produced in the stomach chiefly contain two bases, namely _____ hydroxide and

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Aluminum hydroxide.
Potassium . Calcium
Magnesium ... Chemistry Quiz
Activities Acids, Bases &
Salts ...

Chemistry : Acids, Bases &
Salts - Multiple choice
pH Meter³ solutions of
acids, bases, salts Computer
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research group. A new HTML5
based computer simulation of
this activity is being
developed.

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pH measurement of acid, base and ionic salt solutions ...

A special property of acids and bases is their ability to neutralize the other's properties. In an acid-base (or neutralization)

reaction, the H^+ ions from the acid and the OH^- ions from the base react to create water (H_2O).

Another product of a neutralization reaction is an ionic compound called a salt.

Overview of Acids and Bases - Chemistry LibreTexts

A substance used as an acid-base indicator. (6) 11. An atom or molecule with a positive or negative charge.

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(3) 13. The strength of a base. (10) 15. Common name for sodium hydroxide. (3) 17. Color of litmus paper in acids. (3) Title: Acids, Bases, Salts Crossword Subject: Science Author: Chris Gunn Keywords: Acids and Bases Last modified by ...

Acids, Bases, Salts Crossword

an acid solution < 7 ; a base solution > 7 ; and; a neutral solution $= 7$. (c) The salts of strong acids and weak bases give acidic solution having pH less than 7. Example, NH_4Cl , Ammonium Chloride will have pH less than 7. The salts of weak

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acids and strong bases give basic solution having pH more than 7.

Acids, Bases and Salts Class
10 Important Questions with

...

A _____ is any compound that can be derived from the neutralization of an acid and a base. What is a salt? 100 _____ is a measure of how basic or acidic a solution is. What is the PH scale? 100. Salt water is a _____ base. What is a weak base? 200. If lemon juice has a PH of 2. What type of acid is it? ... If I want to neutralize an acid what do

...

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Answer of Intex -1,

Exercise-3 A ,Exercise-3 B

Exercise-3 C , Miscellaneous

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of Exercise-3 Acids Bases

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In acid – base chemistry,
salts are ionic compounds
that result from the

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neutralization reaction of an acid and a base. Basic salts contain the conjugate base of a weak acid, so when they dissolve in water, they react with water to yield a solution with pH greater than 7.0.

Acid-Base Properties of Salts | Boundless Chemistry Acids/Bases and Salts DRAFT. 10th - 12th grade. 168 times. Chemistry. 78% average accuracy. 7 months ago. emetzger. 0. Save. Edit. Edit. Acids/Bases and Salts DRAFT. ... Many cleaning solutions are bases. Which of the following is a common property among basic

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solutions? answer choices .
slippery feel. white color.
salty taste. sour taste.

Acids/Bases and Salts |
Chemistry Quiz - Quizizz
Equations of Acids, Bases
and Salts: Acid + Metal →
Salt + Hydrogen gas H_2SO_4
+ Zn → $ZnSO_4$ + H_2 ; Base +
Metal → Salt + Hydrogen gas
 $2NaOH$ + Zn → Na_2ZnO_2
(Sodium zincate) + H_2 ; Base
+ Acid → Salt + Water $NaOH$
(aq) + HCl (aq) → $NaCl$ (aq)
+ H_2O (l) Acids give
hydronium ions in water HCl
+ H_2O → H_3O^+ + Cl^-
Bases generate OH^- ions in
water

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10 Notes Science Chapter 2

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