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### Bronsted-Lowry and Lewis

acids and bases. -

Arrhenius: Acid dissolved in water increases conc. Of  $H^+$  ions. Base dissolved in water increases conc. Of  $OH^-$  ions - Bronsted-Lowry: Acid: substance that donates a

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proton to another substance  
Base: substance that accepts  
a ...

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guide, you will find the key concepts of acids and bases and acid-base reactions. It summarizes the factors such as electronegativity, polarizability, induction, and resonance-stabilization for determining the acidity

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of molecules and understating the pKa values and relationship to acid strength. The principle of determining the acid-base equilibrium and main implication of Henderson-Hasselbalch

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equation are also addressed with specific examples.

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- According to the Bronsted-Lowry theory, acids are proton ( $H^+$ ) donors and

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bases are proton acceptors.

- Protons form hydronium ions,  $\text{H}_3\text{O}^+$ , in water when they bond to polar water molecules.
- In the IUPAC system, a carboxylic acid is named by replacing the e ending of the alkane name



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bases pH scale versus pOH scale—what happens to the concentration to  $H^+$  versus  $OH^-$ —as pH goes up and down? Know definitions of acids and bases according to Arrhenius, Bronsted-Lowry, and Lewis. Label Bronsted-

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in water). pH refers to the hydrogen ion concentration, or the ability of a species to donate/accept protons or electrons. Acids and bases reflect the relative availability of hydrogen ions or proton/electron

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of Acids and Alkalis 1.

Strong acids / alkalis are  
corrosive. Weak acids /  
alkalis are generally less



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corrosive. 2. Dilute acids /  
alkalis contain a low  
percentage of acids /  
alkalis. Concentrated acids  
/ alkalis contain a high  
percentage of acids /  
alkalis. 3. Concentrated  
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